**Average Atomic Mass – General Chem**

1. Hydrogen has three isotopes: hydrogen-1, hydrogen-2, and hydrogen-3. If we average the weights of these three isotopes, we find that the average is equal to 2. However, when we look at the periodic table, we find that the average atomic mass is almost exactly 1. How does this happen, given that it is an *average* atomic mass.
2. What does the abundance of each isotope have to do with determining the average atomic mass?
3. If lithium-7 were the only isotope of lithium, would that mean that all atoms of lithium weighed 7 amu? Explain why or why not.
4. Are all isotopes radioactive? Explain why or why not.